



Standard Test Method for Determination of Phosphorus in Iron Ores by Photometric Method¹

This standard is issued under the fixed designation E 1070; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 This test method covers the determination of phosphorus in iron ores, concentrates, and agglomerates in the concentration range from 0.005 to 1.0 % phosphorus.

1.2 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.3 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 *ASTM Standards:*²

D 1193 Specification for Reagent Water

E 50 Practices for Apparatus, Reagents, and Safety Precautions for Chemical Analysis of Metals, Ores, and Related Materials

E 60 Practice for Molecular Absorption Spectrometric Methods for Chemical Analysis of Metals, Ores, and Related Materials

E 135 Terminology Relating to Analytical Chemistry for Metals, Ores, and Related Materials

E 877 Practice for Sampling and Sample Preparation of Iron Ores and Related Materials

E 882 Guide for Accountability and Quality Control in the Chemical Analysis Laboratory

3. Terminology

3.1 *Definitions*—For definitions of terms used in this test method, refer to Terminology **E 135**.

4. Summary of Test Method

4.1 The sample is fused in a zirconium crucible with sodium peroxide. The melt is dissolved in water and hydrochloric acid. In a suitable aliquot, the molybdenum blue complex is formed by the addition of ammonium molybdate-hydrazine sulfate solution. The absorbance of the phospho-molybdenum-blue complex is measured at 725 nm.

5. Significance and Use

5.1 This test method for the analysis of iron ore concentrates and agglomerates is primarily intended as a referee method to test for compliance with compositional specifications. It is assumed that users of this test method will be trained analysts capable of performing common laboratory procedures skillfully and safely. It is expected that work will be performed in a properly-equipped laboratory and that proper waste disposal procedures will be followed. Appropriate quality control practices shall be followed, such as those described in Guide **E 882**.

5.2 The determination of this element is needed for international trade and primary iron and steel making.

6. Interferences

6.1 Elements normally found in iron ores do not interfere excepting arsenic giving positive interference (0.01 % As = 0.001 % P).

¹ This test method is under the jurisdiction of ASTM Committee E01 on Analytical Chemistry for Metals, Ores, and Related Materials and is the direct responsibility of Subcommittee E01.02 on Ores, Concentrates, and Related Metallurgical Materials.

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² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

7. Apparatus

7.1 *Zirconium Crucible*, 50 mL capacity.

7.2 *Spectrophotometer*—Visible spectrophotometer capable of measuring absorbance at the 725 nm wavelength using a 1-cm path length cell in accordance with Practice E 60.

8. Reagents and Materials

8.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.³ Other grades may be used, provided it is first ascertained that the reagent is of sufficient high purity to permit its use without lessening the accuracy of the determination.

8.2 *Purity of Water*—Unless otherwise indicated, references to water shall be understood to mean reagent water as defined by Type II of Specification D 1193.

8.3 *Ammonium Molybdate Solution (20 g/L)*—To 500 mL of water, add cautiously and slowly 300 mL of H₂SO₄ and cool. Add 20 g of ammonium molybdate (NH₄)₆-Mo₇O₂₄·4H₂O. Stir to dissolve and dilute to 1 L with water.

8.4 *Fusion Blank*—Dissolve 4 g of Na₂O₂ in 40 mL of water in a 250-mL beaker. Add 30 mL of HCl. Boil for 2 min. Cool and dilute to 100 mL with water in a volumetric flask. Prepare fresh as needed.

8.5 *Hydrazine Sulfate (1.5 g/L)*—Dissolve 0.15 g of hydrazine sulfate NH₂·NH₂·H₂SO₄ in water and dilute to 100 mL with water. Prepare fresh as needed.

8.6 *Molybdate (5 g/L) Hydrazine Sulfate (0.15 g/L) Solution*—Add 50 mL of ammonium molybdate solution (20 g/L) to 100 mL of water. Add 20 mL of hydrazine sulfate solution (1.5 g/L) and dilute to 200 mL with water.

NOTE 1—This solution should be prepared freshly within 30 min of use in a quantity appropriate for the number of tests being made.

8.7 *Sodium Peroxide (Na₂O₂)*—Use caution when using peroxide.

8.8 *Sodium Sulfite Solution (100 g/L)*—Dissolve 10 g of sodium sulfite (Na₂SO₃) in water and dilute to 100 mL with water.

³ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

8.9 *Standard Phosphorus Solution*—Dry anhydrous disodium phosphate (Na₂HPO₄) at 105°C for 2 h and after desiccation, dissolve 0.2292 g of the reagent in 200 mL of water. Dilute to 1 L with water in a volumetric flask and mix. This solution, A, provides 1 mL = 50 µg P. Transfer 10.00 mL of solution A into a 50-mL volumetric flask, dilute to mark with water, and mix. This solution, B, provides the standard phosphorus solution for calibration 1 mL = 10 µg P.

9. Hazards

9.1 For precautions to be observed in this method, refer to Practices E 50.

10. Sampling and Sample Preparation

10.1 Collect and prepare gross samples in accordance with Practice E 877.

10.2 Pulverize the laboratory sample to pass a No. 100 (150-µm) sieve.

NOTE 2—To facilitate decomposition, some ores, such as specular hematite, require grinding to pass a No. 200 (75-µm) sieve.

11. Calibration and Standardization

11.1 The recommended concentration range is from 0.005 to 0.10 mg phosphorus in 50 mL of color solution using cell depth of 1 cm.

11.2 Into a series of five 150-mL beakers, transfer 10.0 mL of fusion blank solution (8.4) and then transfer 0, 0.50, 2.50 mL of standard B; 1.00 and 2.00 mL of standard A phosphorus solution (8.9) corresponding to 0, 5, 25, 50, and 100 µg of phosphorus respectively. To each beaker add 15 mL of sodium sulfite solution (8.8). Mix. Bring to a boil. Add 20 mL of molybdate-hydrazine sulfate solution (8.6). Bring to a boil. Simmer for 10 min in boiling water bath. Cool. Transfer into a 50-mL volumetric flask and dilute to the mark with water. Mix. This solution is stable for at least 2 h.

11.3 *Photometry*—Adjust the spectrophotometer to the initial setting, using water as the reference solution. While maintaining this setting, measure absorbance of the calibration standard and sample solutions at approximately 725 nm in a 1-cm cell.

11.4 *Preparation of Calibration Curve*—Subtract the absorbance of the 0-mL phosphorus standard solution from the absorbance of each standard solution and plot the net absorbance against µg of phosphorus in the color solution.

12. Procedure

12.1 Weigh approximately the amount of the prepared sample specified in the following table into a small weighing bottle previously dried at 150°C: